MINISTRY OF EDUCATION, HERITAGE AND ARTS

YEAR 13 CHEMISTRY

REVISION WORKSHEET 9

Write the answers to the following questions in your exercise/activity books.

Strand 3: Reactions	**Sub-strand: Electrochemistry**

1. A cell notation is a shorthand way of representing a galvanic cell.

   For the cell notation given below, label the:

   $Sn_{(s)} / Sn^{2+}_{(aq)} // Ag^{+}_{(aq)} / Ag_{(s)}$

   A – Cathode
   B – Anode
   C – Reduction half-cell
   D – Oxidation half-cell
   E – Salt bridge
   F – Phase change/boundary

   (6 marks)

2. Determine the **cell notation** for the following redox reaction.

   $Mn_{(s)} + Cu^{2+}_{(aq)} \longrightarrow Mn^{2+}_{(aq)} + Cu_{(s)}$

   (2 marks)

3. Obtain the overall balanced equation in an **acidic medium** for the following cell notation.

   $Fe_{(s)} / Fe^{2+}_{(aq)} // NO_3^{-}{}_{(aq)} / NO_{(g)} / Pt_{(s)}$

   (6 marks)

The End